

Heat and Material Balances for Nonautogenous Wastes

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INTRODUCTION

The incineration of nonautogenous (will not sustain combustion) wastes is becoming increasingly prevalent as more stringent water quality regulations are being enacted and enforced and land available for waste disposal is becoming increasingly difficult to find.

An understanding of the parameters affecting combustion temperatures and auxiliary fuel usage will aid both the designer and user of incinerators in choosing those items, which collectively form the incinerator system, that will give optimum performance at minimum cost.

The purpose of this paper is to provide the reader with a method of making heat and material balances on nonautogenous wastes. The method of solution is applicable where:

- 1) Auxiliary fuel is mixed directly with the waste in the combustion chamber.
- 2) Auxiliary fuel is supplied by a fuel burner preheating the combustion air.
- 3) Auxiliary fuel is supplied by a fuel burner serving as an after-burner for the combustion products from the incinerator.

The solutions obtained only approximate a real system. Items such as heat losses to surroundings, water vapor present in the air, heat of compression in combustion air, and others have not been included in the calculations. The values to be placed on these

items are usually determined by the specific type of equipment used to perform the operation and judgment of the designer. However, the results of these calculations do enable the designer to graphically display them in such a manner as to indicate the optimum mode of operation.

These calculations have been carried out past the usual significant figures in order to facilitate bookkeeping when making the heat and material balances.

TERMINOLOGY

The following terminology is generally used in describing the physical characteristics of waste sludges.

1. Total Solids (TS) – the percentage, by weight, of the solid fraction of the waste sludge.

NOTE: Unless specifically stated, the liquid fraction is assumed to be water.

2. Volatile Solids (VS) – the percentage, by weight, of the combustible fraction of the Total Solids (TS). Includes fixed carbon. The remaining solids are assumed to be ash.

3. Heating Value (HHV) – can be expressed either as:

- (a) Btu's/lb. TS
- (b) Btu's/lb. VS
- (c) Btu's/lb. Sludge (Total Solids plus water)

NOTE: Bomb Calorimeter results are usually reported as Btu's/lb. TS

Ultimate Analysis of Volatile Solids

Carbon	50%
Hydrogen	7%
Oxygen	39%
Nitrogen	4%
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	100%

CALCULATIONS

For the example calculation presented here, we shall use 100 lb. Total Solids as a calculation base. We shall also assume that the waste sludge has the following analysis:

Total Solids	22% TS
Volatile Solids	80% VS

Btu's/lb. VS 9,500

Let us also assume that we will incinerate the sludge with only 30% Excess Air to minimize auxiliary fuel usage and burn at 1650°F. to insure complete deodorization.

From the preceding we can now determine the feed to the incinerator.

$$100 \text{ lb. TS} \times \frac{80 \text{ lb. VS}}{100 \text{ lb. TS}} = 80 \text{ lb. VS}$$

$$50\% \text{ C} \times 80 \text{ lb. VS} = 40.00 \text{ lb. Carbon}$$

$$7\% \text{ H} \times 80 \text{ lb. VS} = 5.60 \text{ lb. Hydrogen}$$

$$39\% \text{ O} \times 80 \text{ lb. VS} = 31.20 \text{ lb. Oxygen}$$

$$4\% \text{ N} \times 80 \text{ lb. VS} = \underline{3.20 \text{ lb. Nitrogen}}$$

80.00 lb. VS 80.00 lb. VS

$$100 \text{ lb. TS} \times \frac{(100 - 22) \text{ lb. H}_2\text{O}}{22 \text{ lb. TS}} = 354.50 \text{ lb. H}_2\text{O}$$

$$100 \text{ lb. TS} \times \frac{(100 - 80) \text{ lb. Ash}}{100 \text{ lb. TS}} = \frac{20.00 \text{ lb. Ash}}{454.50 \text{ lb. Sludge}}$$

COMBUSTION OF VOLATILE SOLIDS

$$40.00 \text{ lb. C} \times \frac{32 \text{ lb. O}}{12 \text{ lb. C}} = 106.67 \text{ lb. Oxygen}$$

$$5.60 \text{ lb. H} \times \frac{16 \text{ lb. O}}{2 \text{ lb. H}} = \frac{44.80}{151.47}$$

Oxygen available in VS -31.20

0.30 x 120.27 36.08

156.35 x $\frac{76.85 \text{ lb. N}}{23.15 \text{ lb. O}}$ 519.03

675.38 lb. Combustion Air

Combustion Products

$$40.00 \text{ lb. C} \times \frac{44 \text{ lb. CO}_2}{12 \text{ lb. C}} = 146.67 \text{ lb. CO}_2$$

$$5.60 \text{ lb. H} \times \frac{18 \text{ lb. H}_2\text{O}}{2 \text{ lb. H}} = 50.40 \text{ lb. H}_2\text{O}$$

Nitrogen in air	519.03 lb.				
Nitrogen in VS	3.20				
	<u>522.23 lb.</u>	=		522.23 lb. N ₂	
Excess Oxygen				<u>36.08 lb. O₂</u>	
				755.38 lb. Flue Gas	

Calculation Check

Input

Volatile Solids	80.00 lb.				
Combustion Air	<u>675.38 lb.</u>				
	755.38 lb.				

Output

Flue Gas	755.38 lb.				
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COMBUSTION OF AUXILIARY FUEL

Natural Gas:

76% Carbon	0.044/lb./ft. ³				
24% Hydrogen	1,003 Btu's/ft. ³				
22,800 Btu's/lb.					

Combustion of 1 lb. Natural Gas at 30% XS Air.

0.76 lb. C x $\frac{32 \text{ lb. O}}{12 \text{ lb. C}}$	=			2.03 lb. Oxygen	
0.24 lb. H x $\frac{16 \text{ lb. O}}{2 \text{ lb. H}}$	=			<u>1.92 lb.</u>	
	=			3.95 lb. Stoichiometric Oxygen	
0.30 x 3.95				<u>1.19 lb. XS Oxygen</u>	
				5.14 lb. Combustion Oxygen	
5.14 lb. O x $\frac{76.85 \text{ lb. N}}{23.15 \text{ lb. O}}$	=			<u>17.06 lb. Nitrogen</u>	
				22.20 lb. Combustion Air	

Combustion Products

0.76 lb. C x $\frac{44 \text{ lb. CO}_2}{12 \text{ lb. C}}$	=			2.79 lb. CO ₂	
0.24 lb. H x $\frac{18 \text{ lb. H}_2\text{O}}{2 \text{ lb. H}}$	=			2.16 lb. H ₂ O	
Excess Oxygen				1.19 lb. O ₂	
Nitrogen in combustion air				<u>17.06 lb. N₂</u>	
				23.20 lb. Flue Gas	

Calculation Check

Input

1.00 lb. Natural Gas
22.20 lb. Combustion Air
 23.20 lb.

Output

23.20 lb. Flue Gas

HEAT AND MATERIAL BALANCE

Enthalpy of combustion products at 1650°F.

Carbon Dioxide	CO ₂	419 Btu's/lb.
Water	H ₂ O	1839
(Including Heat of Vaporization)		
Nitrogen	N ₂	417
Oxygen	O ₂	386
Ash	Ash	413

Btu Output from Sludge (Volatile Solids, Water, and Ash) @ 1650°F.

146.67 lb. CO ₂	x	419 Btu's/lb.	=	61,455 Btu's
50.40 lb. H ₂ O	x	1839	=	92,686
522.23 lb. N ₂	x	417	=	217,777
36.08 lb. O ₂	x	386	=	13,927
354.50 lb. H ₂ O	x	1839	=	651,926
20.00 lb. Ash	x	413	=	8,260
				<u>1,046,031 Btu's</u>

Btu Output from 1 lb. Natural Gas @ 1650°F.

2.79 lb. CO ₂	x	419	=	1,169 Btu's
2.16 lb. H ₂ O	x	1839	=	3,972
17.06 lb. N ₂	x	417	=	7,114
1.19 lb. O ₂	x	386	=	459
				<u>12,714 Btu's</u>

Determination of Natural Gas requirements

Let X equal the pounds of Natural Gas required

Input

80.00 lb. VS x 9,500 Btu's/lb. VS = 760,000 Btu's
 X lb. Natural Gas x 22,800 Btu's/lb. = 22,800X Btu's

Output

Sludge 1,046,031 Btu's
 Natural Gas 12,714X

Input	=	Output
760,000 + 22,800X	=	1,046,031 + 12,714X
10,086X	=	286,031
X	=	28.36 lb. Natural Gas

Total Air Requirements

For combustion of VS 675.38 lb.

For combustion of Natural Gas

$$28.36 \text{ lb. gas} \times \frac{22.20 \text{ lb. air}}{1.00 \text{ lb. gas}} = \frac{629.59}{1,304.97 \text{ lb. air}}$$

Combustion Products from Natural Gas

$$28.36 \text{ lb. gas} \times 2.79 \text{ lb. CO}_2/\text{lb. gas} = 79.12 \text{ lb. CO}_2$$

$$28.36 \text{ lb. gas} \times 2.16 \text{ lb. H}_2\text{O}/\text{lb. gas} = 61.26 \text{ lb. H}_2\text{O}$$

$$28.36 \text{ lb. gas} \times 1.19 \text{ lb. O}_2/\text{lb. gas} = 33.75 \text{ lb. O}_2$$

$$28.36 \text{ lb. gas} \times 17.06 \text{ lb. N}_2/\text{lb. gas} = 483.82 \text{ lb. N}_2$$

Total Weight of Combustion Products (Sludge + Natural Gas)

225.79 lb. CO₂

466.16 lb. H₂O

69.83 lb. O₂

1,006.05 lb. N₂

1,767.83 lb. Flue Gas/100 lb. TS

Calculation Check

Input

Air	1,304.97 lb.
Sludge	454.50 lb.
Natural Gas	<u>28.36 lb.</u>
	1,787.83 lb.

Output

Flue Gas	1,767.83 lb.
Ash	<u>20.00 lb.</u>
	1,787.83 lb.

FLUE GAS VOLUMES

$$V = \frac{mRT}{p}$$

$$V = \text{ft.}^3$$

$$P = \text{lb.}/\text{ft.}^2$$

$$m = \text{lbm.}$$

$$R = \text{ft.}\cdot\text{lb.}/\text{lbm.}\cdot^\circ\text{R}$$

$$T = ^\circ\text{R}$$

$$\text{CO}_2 \quad R = 35.10 \text{ ft.}\cdot\text{lb.}/\text{lbm.}\cdot^\circ\text{R}$$

$$\text{H}_2\text{O} \quad R = 85.76$$

$$\text{N}_2 \quad R = 55.15$$

$$\text{O}_2 \quad R = 48.28$$

For the example calculation:

$$T = 1650 + 460 = 2110^\circ\text{R}$$

$$P = 14.7 \text{ lb.}/\text{in.}^2 \times 144 \text{ in.}^2/\text{ft.}^2 = 2116.8 \text{ lb.}/\text{ft.}^2$$

$$V = \frac{2110.0}{2116.8} \times m \times R = 0.99679 \times m \times R$$

CO ₂	0.99679	x	225.79	x	35.10	=	7,900 ft. ³	=	7.42%
N ₂	0.99679	x	1,006.05	x	55.15	=	55,306	=	51.97%
O ₂	0.99679	x	69.83	x	48.28	=	3,361	=	3.16%
H ₂ O	0.99679	x	466.16	x	85.76	=	39,850	=	37.45%
							<u>106,417 ft.³</u>		<u>100.00%</u>

On a dry basis (Orsat Analysis)

CO ₂					7,900 ft. ³	=	11.87%
N ₂					55,306	=	83.08%
O ₂					3,361	=	5.05%
					<u>66,567 ft.³</u>		<u>100.00%</u>

DISCUSSION

In order to cover a wide range of input conditions, a computer program was developed to perform the calculations described in this paper. Listed below are the parameters used in these calculations.

Btu/lb. VS	6,350; 9,500; 12,500
% TS	10; 15; 20; 25; 30; 35
% VS	60; 70; 80

% XS Air 40; 60; 80; 100; 120

Temperature °F. 1,400; 1,500; 1,600; 1,700; 1,800

Thirteen hundred and fifty (1,350) separate heat and material balances were made. Several graphs showing the results of these calculations are included in the appendix of this paper. With the aid of graphs such as these, the designer can evaluate alternatives to a particular incineration problem.

ENTHALPY OF FLUE GASES ABOVE 80°F

RANGE 1400°F. to 1800°F.

CO₂ h = 0.30325 T - 81.1 Btu/lb.

N₂ h = 0.28550 T - 54.3 Btu/lb.

O₂ h = 0.26478 T - 50.7 Btu/lb.

H₂O h = 0.57133 T + 898.7 Btu/lb.

(including 1048.6 Btu/lb.
heat of vaporization)

Ash h = 0.25 T - 20 Btu/lb.

T = Temperature °F.

ANALYSIS OF VOLATILE SOLIDS

1. 12,500 Btu/lb. VS

65.0% Carbon

9.0% Hydrogen

21.0% Oxygen

5.0% Nitrogen

2. 9,500 Btu/lb. VS

50.0% Carbon

7.0% Hydrogen

39.0% Oxygen

4.0% Nitrogen

3. 6,350 Btu/lb. VS

38.2% Carbon

7.3% Hydrogen

54.4% Oxygen

0.1% Nitrogen

APPENDIX



